

Chemistry Stoichiometry Quiz 1 Answers

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~~Balance $\text{Ca}(\text{OH})_2(\text{aq}) + \text{H}_3\text{PO}_4(\text{aq}) \rightarrow \text{Ca}_3(\text{PO}_4)_2(\text{s}) + \text{H}_2\text{O}(\text{l})$ (s): There is 1 mole of water per mole calcium phosphate. There are 2 moles of water produced per mole of calcium phosphate. There are 6 moles of water produced per mole of calcium phosphate. There are 10 moles of water produced per mole of calcium phosphate. 8.~~

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~~$\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$ 4. Consider the following balanced equation. $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 3\text{O}_2 \rightarrow 12\text{CO}_2 + 11\text{H}_2\text{O}$ Determine the mass of citric acid ($\text{C}_6\text{H}_8\text{O}_7$) produced when 2.5 mol $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ is used. 5. Chloroform (CHCl_3) is produced by a reaction between methane and chlorine.~~

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Question 1. SURVEY. 30 seconds. Q. The number of significant figures in an answer to a stoichiometry problem is determined only by. answer choices. the number of significant figures of any measured quantities in the problem. the number of decimal places in the molar masses of substances in the chemical equation.

Balance Ca (OH) (aq) + H₃PO₄ (aq) → H₂O (g) + Ca₃(PO₄)₂ (s): There is 1 mole of water per mole calcium phosphate. There are 2 moles of water produced per mole of calcium phosphate. There are 6 moles of water produced per mole of calcium phosphate. There are 10 moles of water produced per mole of calcium phosphate. 8.

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