

Chapter 18 Reaction Rates Equilibrium Worksheet Answers

Chapter 18: Reaction Rates & Equilibrium. If a stress is applied to a system in dynamic equilibrium, the system changes in a way that relieves the stress.

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Chapter 18 - Reaction Rates & Equilibrium This chapter examines the idea of reversible reactions and their equilibrium positions. Significant emphasis is placed on how equilibrium and the rate of a reaction can be affected by altering temperature, pressure and concentrations.

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Figure 18.2, page 542: compare the rates A “rate” is a measure of the speed of any change that occurs within an interval of time In chemistry, reaction rate is expressed as the amount of reactant changing per unit time. Example: 3 moles/year, or 5 grams/second

Chapter 18 “Reaction Rates and Equilibrium”

Chapter 18 Reaction Rates And Equilibrium. _____ the temperature of a chemical reaction causes the equilibrium position of a reaction to shift in the direction that absorbs heat.

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The specific rate constant (k) for a reaction is a proportionality constant relating the concentrations of reactants to the rate of the reaction. $aA + bB \rightarrow cC + dD$ For the reaction of A with B, the rate of reaction is dependent on the concentrations of both A and B.

Chapter 18: Reaction Rates and Equilibrium

a reaction in which the conversion of reactants into products and the conversion of products into reactants occur simultaneously chemical equilibrium a state of balance in which the rates of the forward and reverse reactions are equal; no net change

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Chapter 18 Notes Reaction Rates and Equilibrium 18.1 Rates of Reaction Collision Theory
o Rate = The speed of any change that occurs within an interval of time
o KEY = In chemistry, the rate of chemical change or the reaction rate is usually expressed as the amount of reactant changing per unit time

Chapter 18 Notes Reaction Rates and Equilibrium

first-order reaction: a reaction in which the rate is directly proportional to the concentration of one of the reactants
reaction rate: the number of particles that react in a given time to form products
Le Chatelier's principle: If a stress is applied to a system in dynamic equilibrium, the system changes to relieve the stress
elementary reaction

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RATES OF REACTION
Speeds of chemical reactions can be extremely fast or extremely slow. The rate is a measure of the speed of any change that occurs within an interval of time (which could range from fractions of a second to centuries).

Chapter 18 Reaction Rates and Equilibrium

Chapter 18 Reaction Rates and Equilibrium Part 2 3/13/11. Chapter 18 Reaction Rates and Equilibrium Part 2 3/13/11. ... 18. What is a specific ... The constant is a proportionality constant relating the concentrations of chemical change to the rate of the reaction. B.

Chapter 18 Reaction Rates And Equilibrium Part 2 3/13/11 ...

Chapter 18 - Reaction Rates and Equilibrium - 18.1 Rates of Reaction - 18.1 Lesson Check - Page 601: 2
Answer The rate of a chemical reaction is dependent on temperature, concentration, particle size, and the use of a catalyst.

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