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average atomic  
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adding fraction-of-  
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for the three  
isotopes—i.e., using  
Equation 2.2.

(Remember:

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fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus. A Si =  $f_{28} A_{28} + f_{29} A_{29} + f_{30} A_{30}$  Si Si Si Si Si

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Edition AUTHOR:

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reference for  
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associated with  
the Bohr model  
of the atom are  
(1) that  
electrons are

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particles  
moving in  
discrete  
orbitals, and  
(2) electron  
energy is  
quantized into  
shells.

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Solution The average atomic weight of silicon ( A ) is computed by adding fraction-of-occurrence/atomic weight Si products for the three isotopes—i.e., using Equation 2.2.

(Remember: fraction

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of occurrence is equal to the percent of occurrence divided by 100.) Thus.  $A_{Si} = \frac{f_{28}}{28 + f_{29}}$   
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