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Chemistry 2 – Atomic Structure. II. Atomic Scientists – How has the theory of the atom changed over time? III. Bohr Models. ... Answer the following questions completely: 1. ... What is the meaning of the word atom? I-G. Chapter 4–2 Guided Reading. HOW SMALL IS AN ATOM? WHAT ' S INSIDE AN ATOM? HOW DO ATOMS OF DIFFERENT ELEMENTS DIFFER ...

Atomic Number Atomic number (Z) of an element is the number of protons in the nucleus of each atom of that element. 79 79 Gold 15 15 Phosphorus 6 6 Carbon Atomic # (Z) # of protons Element 29. Mass Number Mass number is the number of protons and neutrons in the nucleus of an isotope: Mass # = p + + n 0 8 8 18 18 Arsenic 75 33 75 Phosphorus 15 31 16 15 - 31 42 33 - 10 Oxygen - Mass # e - n 0 p ...

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Guided Notes: Atomic Structure 01 Gold Foil Experiment Sketch and label the Rutherford Gold Foil Experiment. In this model, where are each of the subatomic particles? o Proton – o Neutron – Distinguishing among Elements Elements differ in their number of _____.

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Name: Period: Unit 3 Atomic Structure- Guided Notes

Atomic Structure MODULE - 2 Notes Atomic Structure and Chemical Bonding chemistry has been defined as the study of matter in terms of its structure, composition and the properties. As you are aware, matter is made up of atoms, and therefore an understanding of the structure of atom is very important.

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